

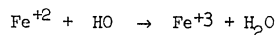
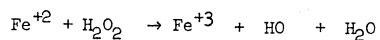
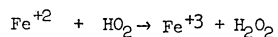
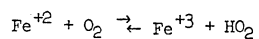
TABLE 3. SOLUBILITY OF $\text{Fe}(\text{OH})_2$ AND $\text{Fe}(\text{OH})_3$ FROM K_{sol} AT 25 C

$K_{\text{sol}} \text{Fe}(\text{OH})_2 = 5.25 \times 10^{-14}$; $K_{\text{sol}} \text{Fe}(\text{OH})_3 = 1 \times 10^{-38}$ (Also 10^{-40} is reported.)

pH	Fe^{+2} , ppm	Fe^{+3} , ppm
1		56000
2		560
3		0.56
4		0.00056
5		
6	29400	
7	294	
8	2.94	
9	0.0294	

b. Oxidation of Ferrous Iron - Weiss studied the mechanism of the ferrous

oxidation and concluded that the reaction proceeded according to the following mechanism:



The above reactions are not balanced with respect to H^+ and, therefore, do not show the possible dependence of reaction rate on this variable.

Huffman and Davidson, 1956 (23) found that "the reaction

